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cations and anions. (When metal gives electron ...

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Electronegativity greater than 1.7 and is a transfer of electrons.

The Lower Electronegativity transfers to the higher electronegativity) t. Ionic Bond. Ions pack in a crystal lattice structure. Lattice Energy. Cation {+} and anion {-} create _____

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6 REVIEW Chemical Bonding SECTION 4 SHORT ANSWER Answer

the following questions in the space provided. 1. b In metals, the

valence electrons are considered to be (a) attached to particular

positive ions. (c) immobile

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chemical bonding that results from the electrical attraction
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1. Use the concept of potential energy to describe how a covalent bond forms between two atoms. As the atoms involved in the formation of a covalent bond approach each other, the

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6 Chemical Bonding

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Chemical Bonding CHAPTER 6 Chemical Bonding SECTION 1
Introduction to Chemical Bonding OBJECTIVES 1. Define Chemical bond. 2. Explain why most atoms form chemical bonds. 3. Describe ionic and covalent bonding.. 4.

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SECTION 3 SHORT ANSWER Answer the following questions in the space provided 1 ____ The notation for sodium chloride, NaCl, stands for one (a) formula unit (c) crystal (b) molecule (d) atom 2 ____ In a crystal of an ionic compound, each ...

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Chapter 6 Review

CHAPTER 12 REVIEW Solutions SECTION 2 SHORT ANSWER

Answer the following questions in the space provided. 1. The following are statements about the dissolving process. Explain each one at the molecular level. a. Increasing the pressure of a solute gas above a liquid solution increases the solubility of the gas

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