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Bond. A mutual
electrical attraction
between the nuclei and
valence electrons of
different atoms that

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binds the atoms
together. (Valence
cloud) Ionic Bonding

cloud) Ionic Bond.
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results from the
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anions. (When metal
gives electron ...

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STUDY. PLAY.
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Electronegativity

greater than 1.7 and is a transfer of electrons.

The Lower

Electronegativity

transfers to the higher electronegativity) t.

Ionic Bond. Ions pack in a crystal lattice

structure. Lattice

Energy. Cation {+}

and anion {-} create

—

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Chemical Bonding

SECTION 4 SHORT

ANSWER Answer the
following questions in
the space provided. 1.

b In metals, the
valence electrons are
considered to be (a)
attached to particular

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positive ions. (c)
immobile.

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chemical bonding that
results from the
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between large
numbers of cations
(metals) and anions
(nonmetals) ... Modern
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Vocabulary. 20 terms.

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7. c 8. a 9. lanthanides
10. 2 11. fourth 12.
transition elements 13.

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32 14. valence
electrons 15. electron
affinity 16.
electronegativity 17.
ionization energy 18. 3
s² 3p⁴ 19 ...

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15. 191 terms.

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Geometries. 20 terms.
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Chemical Bonding
SECTION 2 SHORT
ANSWER Answer the
following questions in
the space provided. 1.
Use the concept of

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potential energy to describe how a covalent bond forms between two atoms. As the atoms involved in the formation of a covalent bond approach each other, the

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CHAPTER 6 Chemical
Bonding SECTION 1

Introduction to
Chemical Bonding
OBJECTIVES 1. Define
Chemical bond. 2.
Explain why most
atoms form chemical
bonds. 3. Describe
ionic and covalent
bonding.. 4.

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SECTION 3 SHORT

ANSWER Answer the following questions in the space provided 1

_____ The notation for sodium chloride, NaCl, stands for one (a) formula unit (c) crystal (b) molecule (d) atom 2

_____ In a crystal of an ionic compound, each

...

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Frequency and Energy

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Solutions SECTION 2
SHORT ANSWER

Answer the following questions in the space provided. 1. The following are statements about the dissolving process. Explain each one at the molecular level. a. Increasing the pressure of a solute gas above a liquid solution increases the solubility of the gas

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