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CHAPTER 5 What You'll
Learn You will compare
the wave and particle
models of light. You

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will describe how the frequency of light emitted by an atom is a unique characteristic of that atom. You will compare and contrast the Bohr and quantum mechanical models of the atom. You will express the arrangements of electrons in atoms through orbital

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Jennie L. Borders.

Section 5.1 - Models of
the Atom. The

Rutherford's model of
the atom did not

explain how an atom

can emit light or the

chemical properties of

an atom. Plum Pudding

Model Rutherford's

Model. The Bohr Model.

Chapter 5 -
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Electrons in Atoms

...are the way electrons are arranged in various orbitals around the nuclei of atoms. Three rules tell us how:

Aufbau principle - electrons enter the lowest energy first.

This causes difficulties because of the overlap of orbitals of different energies - follow the diagram! Pauli

Exclusion Principle - at most 2 electrons per orbital

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Chapter 5 - electrons in
atoms (handouts)

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table & trends

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bonding & chemical
names & formulas

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Science / Chapter 5 - electrons in atoms (handouts)

Chapter 5: Electrons in
Atoms Models of the
Atom Rutherford used
existing ideas about
the atom and proposed
an atomic model in
which the electrons
move around the

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nucleus, like the planets move around the sun. Rutherford's model fails to explain why objects change color when heated.

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Cristina Chemistry I H.
Terms in this set (82)
Rutherford's model-
also called nuclear
model-did not explain
how the electrons were
arranged

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in Atoms. Chapter 5 -
Electrons in Atoms.
Section 5.1 - Models of
the Atom. The
Rutherford's model of

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the atom did not explain how an atom can emit light or the chemical properties of an atom. Plum Pudding Model Rutherford's Model. The Bohr Model.

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Electrons in Atoms

Section 55.1.1 Figure

5.1 Different elements
can have similar
reactions with water.

Objectives Compare
the wave and particle
natures of light. Define
a quantum of energy,
and explain how it is
related to an energy
change of matter.

Contrast continuous
electromagnetic

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IB Chapter Questions.
02:25. Problem 1

Objects get their colors
from reflecting only
certain wavelengths
when hit with white
light. Light reflected
from a green leaf is
found to have a
wavelength of 4.90×10^{-7}
 m . What
is the frequency of the
light? ...

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How are electrons
arranged in an atom?
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Chapter 5“Electrons in Atoms ... Describe the energies and positions of electrons according to the quantum mechanical model.

Section 5.1Models of the Atom. OBJECTIVES: Describe how the shapes of orbitals related to different sublevels differ. Ernest Rutherford's Model.

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can each p orbital

hold? Chapter 5:

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